





If 120 g of carbon reacts with 160 g of  $\text{Br}_2$  and the reaction goes in 50% yield, how many grams of  $\text{Br}_2\text{C}=\text{CBr}_2$  ( $\text{C}_2\text{Br}_4$ ) are produced? Clearly show and explain all your work.

5. (10 pts) Sketch out the Lewis dot structures for the following molecules (use lines for bonds and pairs of dots for lone pairs). Indicate the formal charges, if any, on the appropriate atoms.

